# **Inorganic Chemistry**

# A "Strongly" Self-Catenated Metal–Organic Framework with the Highest Topological Density among 3,4-Coordinated Nets

Huiqing Ma,<sup>†</sup> Di Sun,<sup>†</sup> Liangliang Zhang,<sup>‡</sup> Rongming Wang,<sup>‡</sup> Vladislav A. Blatov,<sup>\*,§,||</sup> Jie Guo,<sup>†</sup> and Daofeng Sun<sup>\*,†,‡</sup>

<sup>†</sup>Key Lab for Colloid and Interface Chemistry of Education Ministry, School of Chemistry and Chemical Engineering, Shandong University, Jinan 250100, People's Republic of China

<sup>§</sup>Department of Chemistry, Samara State University, Ac. Pavlov St. 1, Samara 443011, Russia

<sup>‡</sup>College of Science, China University of Petroleum (East China), Qingdao Shandong 266555, People's Republic of China

<sup>II</sup>Chemistry Department, Faculty of Science, King Abdulaziz University, Jeddah 21589, Saudi Arabia

**Supporting Information** 

**ABSTRACT:** A new type of 3D "strongly" self-catenated metal–organic framework (**SDU-9**) has been constructed from  $[Cu_2(COO)_4]$  paddlewheel secondary building units and a tripodal carboxylate linker. **SDU-9** ( $[Cu_6(H_2O)_6L_4]$ . 24H<sub>2</sub>O, where  $[H_3L = 4,4',4''-(hydroxysilanetriyl)tris-(triphenyl-4-carboxylic acid), represents a rare example of a highly symmetrical coordination network and extremely tight self-catenation. To the best of our knowledge,$ **SDU-9**has the highest topological density among all known 3,4-coordinated nets.

Metal-organic frameworks (MOFs) have caused an upsurge of interest not only for their potential applications in gas adsorption, catalysis, ion exchange, biomedical applications, etc.,<sup>1-6</sup> but also for their intriguing variety of architectures and topologies.<sup>4a,7</sup> The network topological approach proposed by Wells<sup>8</sup> is an important and essential aspect of the analysis, comparison, and design of MOFs by reducing multidimensional structures to simple node-and-connection reference nets. In the past decade, many structural types and entanglement topological features unprecedented in the world of inorganic compounds and minerals have been unmasked in the MOF field.<sup>9</sup> Recently, Yang and co-workers described a silver coordination polymer with a record 54 interpenetrating networks, which provide a new template for future highly interpenetrated or self-catenated structures.<sup>10</sup>

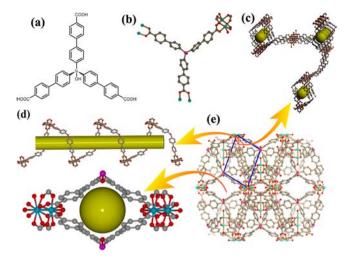
Among the different types of entanglements, such as interpenetration, polycatenation, polythreading, and self-catenation (other equivalent terms are self-penetration or polyknotting), interpenetrating networks are the most abundant and most comprehensively studied objects, as shown in the extensive reviews.<sup>11</sup> However, the study of self-catenated networks considered as extended periodic equivalents of molecular knots remains largely unexplored.<sup>11a</sup> These structures are single networks with regions in which chains from the same net pass through smallest topological circuits in a fashion similar to that of interpenetrating systems. Recently, uninodal (with one kind of node) self-catenated coordination networks were reviewed by Ke et al.<sup>12</sup> According to the list given by Batten in a web site (http:// www.chem.monash.edu.au/staff/sbatten/interpen/examples7. html), a few examples of self-catenated networks with nodes of different connectivity such as 4,4-,<sup>13</sup> 3,5-,<sup>14</sup> 3,6-,<sup>15</sup> and 3,12connections<sup>16</sup> have been observed; however, reports on a selfcatenated 3,4-coordinated network are quite rare. Mu et al. reported an unusual 3,4-coordinated self-catenated network that can be viewed as the cross-linking of a  $2D + 2D \rightarrow 3D$  inclined polycatenation.<sup>17</sup> Similarly, most of the self-catenated networks can be represented as interweaved arrays of 2D or 3D individual equivalent components after breaking some edges,<sup>18</sup> which we call the "weakly" self-catenated nets, in order to discriminate the net reported here. Generally, most reported self-catenated nets are constructed from flexible organic ligands; reports on rigid carboxylate ligands are somewhat rare. To construct such selfcatenated nets, the conformation and linking mode of the organic ligand seem very important.

Recently, we reported a series of porous 3,24-coordinated **rht**type MOFs based on  $C_{3\nu}$ -symmetry silicon-based hexatopic carboxylate linkers.<sup>19</sup> According to the *TOPOS TTD* collection,<sup>11b</sup> in total there are 20 examples of **rht**-type coordination networks, the first of which were synthesized in 2008. Developing our previous results, in this work, we have synthesized a new tripodal organic ligand, 4,4',4''-(hydroxysilanetriyl)tris(triphenyl-4-carboxylic acid) (H<sub>3</sub>L). Surprisingly, the solvothermal assembly of H<sub>3</sub>L and copper paddlewheel secondary building units (SBUs) resulted in the formation of a novel "strongly" self-catenated MOF ([Cu<sub>6</sub>(H<sub>2</sub>O)<sub>6</sub>L<sub>4</sub>]·24H<sub>2</sub>O, which is quite different from other MOFs based on tripodal carboxylate ligands and paddlewheel SBUs.

The solvothermal reaction of  $Cu(NO_3)_2 \cdot 6H_2O$  with  $H_3L$  (Figure 1a) in *N*,*N*-dimethylformamide/ethanol/water [1:1:1 (v/v/v), 10 mL] at 75 °C for 2000 min yields cubic green crystals of **SDU-9**. Single-crystal X-ray diffraction analysis reveals that **SDU-9** crystallizes in the chiral cubic space group *F*432. Two copper ions are bridged by four carboxylates to form the well-known paddlewheel  $[Cu_2(COO)_4]$  SBU (Figure 1b). Each SBU connects four organic ligands, and each ligand binds three SBUs to form a 3D 3,4-coordinated self-catenated MOF containing right-handed helices along the [1,0,1] direction (Figure 1d). The

 Received:
 July 23, 2013

 Published:
 September 23, 2013



**Figure 1.** (a) Organic ligand used in this work. (b) Coordination environment of  $Cu^{II}$  ions in **SDU-9** with thermal ellipsoids at the 50% probability level. Hydrogen atoms were omitted for clarity. (c) Connection type between the helical chains (viewed along the [0,1,-1] direction). (d) Perspective view of the helical chain along the [1,0,1] direction in **SDU-9** (top) and the 1D channel (bottom). (e) 3D packing of **SDU-9** along [1,0,0] direction.

connection type between the helices is different (Figure 1c): one is connected through the  $[Cu_2(COO)_4]$  SBU, and the other is linked through the silicon atom.

To perform topological analysis of the 3D architecture, we have used the *TOPOS* program package.<sup>20</sup> Topologically, by regarding each Cu<sub>2</sub>-SBU as a 4-coordinated node (Figure 2) and

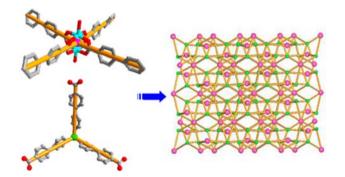
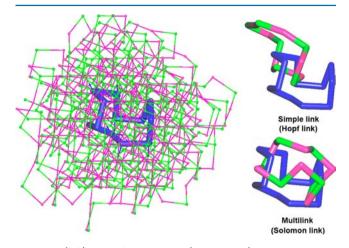


Figure 2. Schematic representation of the 3D framework of SDU-9, where the purple balls are the simplified paddlewheel units.

each L as a 3-coordinated node (Figure 2), the overall 3D net can be rationalized as a binodal 3,4-coordinated edge-transitive (with one kind of edge) underlying net (i.e., the net of centroids of structural groups) with the point symbol<sup>21</sup> of  $(12^3)_4(12^6)_3$ (Figure 2). It is striking that this net has collisions; i.e., some of its vertices have coordinates that are the average of the coordinates of their neighbors.<sup>11c,d</sup> This feature is quite rare for the underlying nets.<sup>11c</sup> Topological classification with the *TOPOS TTD* collection revealed that such a net has never been found in crystals; we have deposited it into the TTD collection under the name **sdf1**. Further analysis showed that there are four kinds of 12-rings in this 3D framework (Figure S4 in the Supporting Information). Each ring consists of six Cu<sub>2</sub>-SBUs that are connected by six halves of organic ligands.

The most fascinating and peculiar structural feature of **SDU-9** is that all of the rings are catenated together, coming to a

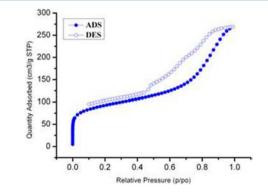
complete framework. Because the underlying net is edgetransitive (all edges are equivalent by symmetry), it cannot be naturally split into separate interpenetrating 2D or 3D nets; after breaking any edge, one obtains an unconnected structure. Hence, we call such self-catenation "strong". The self-catenation is extremely tight: each 12-ring is crossed by more than 100 other 12-rings. Figure 3 demonstrates the unprecedented complexity



**Figure 3.** (left) One of the 12-rings (12b, in blue) catenated by 176 other 12-rings in the underlying net **sdf1** of **SDU-9**. (right) One of the 174 simple links between the 12-rings and one of two multilinks.

of self-catenation: all depicted nodes of the underlying net belong to the 176 12-rings that catenate the same (blue) 12-ring. Besides simple (Hopf-type) links between 12-rings, there are a few multilinks (also known as Solomon links)<sup>22</sup> that are rare in selfcatenated coordination networks (Figure 3). This tight selfcatenation causes a high topological density<sup>23</sup> of the net,  $TD_{10} =$ 3245; according to the RCSR database,<sup>24</sup> to the best of our knowledge, this is the highest topological density among all known 3,4-coordinated nets.

On the basis of the calculation from the *PLATON/VOID* routine,<sup>25</sup> the total solvent-accessible volume for the desolvated framework after removal of guest solvents and coordinated water molecules is estimated to be 35.9%. To check the permanent porosities of **SDU-9**, the freshly prepared samples were soaked in methanol and dichloromethane to exchange the less volatile solvent molecules. A color change from bright-blue to blue-green, indicating that open Cu<sup>II</sup> sites similar to those observed for other frameworks have been generated.<sup>18</sup> As shown in Figure 4, desolvated crystal **SDU-9** displays a typical type IV adsorption



**Figure 4.**  $N_2$  sorption isotherms at 77 K (solid circles, adsorption; open circles, desorption).

#### **Inorganic Chemistry**

In summary, a 3D  $(12^3)_4(12^6)_3$  coordination network with high topological density and extremely tight self-catenation of 12-rings has been obtained based on  $[Cu_2(COO)_4]$  paddlewheel SBUs and a tripodal carboxylate linker. The self-catenation is "strong"; i.e., the network cannot be transformed into an array of interpenetrating nets by breaking any chemical bond.

# ASSOCIATED CONTENT

#### **Supporting Information**

X-ray crystallographic data in CIF format, detailed synthetic procedures, IR and TGA for the compounds, and additional graphics. This material is available free of charge via the Internet at http://pubs.acs.org.

#### AUTHOR INFORMATION

### **Corresponding Authors**

\*E-mail: blatov@samsu.ru.

\*E-mail: dfsun@sdu.edu.cn.

# Funding

This work was supported by the NSFC (Grants 21001115 and 21271117), NCET-11-0309 and the Shandong Natural Science Fund for Distinguished Young Scholars (JQ201003), and the Fundamental Research Funds for the Central Universities (Grants 13CX05010A and 13CX02006A).

#### Notes

The authors declare no competing financial interest.

## REFERENCES

(1) (a) Zhou, H. C.; Long, J. R.; Yaghi, O. M. Chem. Rev. 2012, 112, 673.
 (b) Suh, M. P.; Park, H. J.; Prasad, T. K.; Lim, D.-W. Chem. Rev. 2012, 112, 782.
 (c) Sumida, K.; Rogow, D. L.; Mason, J. A.; McDonald, T. M.; Bloch, E. D.; Herm, Z. R.; Bae, T.-H.; Long, J. R. Chem. Rev. 2012, 112, 724.
 (d) Getman, R. B.; Bae, Y.-S.; Wilmer, C. E.; Snurr, R. Q. Chem. Rev. 2012, 112, 703.

(2) (a) Farha, O. K.; Özgür Yazaydın, A.; Eryazici, I.; Malliakas, C. D.; Hauser, B. G.; Kanatzidis, M. G.; Nguyen, S. T.; Snurr, R. Q.; Hupp, J. T. *Nat. Chem.* **2010**, *2*, 944. (b) Lu, G.; Li, S.; Guo, Z.; Farha, O. K.; Hauser, B. G.; Qi, X.; Wang, Y.; Wang, X.; Han, S.; Liu, X.; DuChene, J. S.; Zhang, H.; Zhang, Q.; Chen, X.; Ma, J.; Loo, S. C.; Wei, W. D.; Yang, Y.; Hupp, J. T.; Huo, F. *Nat. Chem.* **2012**, *4*, 310. (c) McKinlay, R. M.; Cave, G. W. V.; Atwood, J. L. *Proc. Natl. Acad. Sci. U.S.A.* **2005**, *102*, 5944.

(3) (a) Long, J. R.; Yaghi, O. M. Chem. Soc. Rev. 2009, 38, 1213.
(b) Horcajada, P.; Gref, R.; Baati, T.; Allan, P. K.; Maurin, G.; Couvreur, P.; Férey, G.; Morris, R. E.; Serre, C. Chem. Rev. 2012, 112, 1232.
(c) Corma, A.; Garcia, H.; Xamena, F. X. L. Chem. Rev. 2010, 110, 4606.
(d) Cui, Y. J.; Yue, Y. F.; Qian, G. D.; Chen, B. L. Chem. Rev. 2012, 112, 1126. (e) Yoon, M.; Srirambalaji, R.; Kim, K. Chem. Rev. 2012, 112, 1196.

(4) (a) Yaghi, O. M.; O'Keeffe, M.; Ockwig, N. W.; Chae, H. K.; Eddaoudi, M.; Kim, J. *Nature* **2003**, *423*, 705. (b) Müller, U.; Schubert, M.; Teich, F.; Puetter, H.; Schierle-Arndt, K.; Pastre, J. J. *Mater. Chem.* **2006**, *16*, 626. (c) Dinca, M.; Long, J. R. *Angew. Chem., Int. Ed.* **2008**, *47*, 6766. (d) Czaja, A. U.; Trukhan, N.; Müller, U. *Chem. Soc. Rev.* **2009**, *38*, 1284.

(5) (a) Ma, L.; Abney, C.; Lin, W. Chem. Soc. Rev. 2009, 38, 1248.
(b) Czaja, A. U.; Trukhan, N.; Müller, U. Chem. Soc. Rev. 2009, 38, 1284.
(c) Allendorf, M. D.; Bauer, C. A.; Bhakta, R. K.; Houk, R. J. T. Chem. Soc. Rev. 2009, 38, 1330.

(6) (a) Blake, A. J.; Champness, N. R.; Hubberstey, P.; Li, W. S.; Withersby, M. A.; Schröder, M. *Coord. Chem. Rev.* **1999**, *183*, 117. (b) Moulton, B.; Zaworotko, M. J. Chem. Rev. 2001, 101, 1629.
(c) Hagrman, P. J.; Hagrman, D.; Zubieta, J. Angew. Chem., Int. Ed. 1999, 38, 2638.

(7) Hill, R. J.; Long, D. L.; Champness, N. R.; Hubberstey, P.; Schröder, M. Acc. Chem. Res. 2005, 38, 335.

(8) Wells, A. F. Three-dimensional Nets and Polyhedra; Wiley-Interscience, New York, 1977.

(9) (a) Batten, S. R.; Robson, R. Angew. Chem., Int. Ed. 1998, 37, 1460.
(b) Batten, S. R. CrystEngComm 2001, 3, 67.

(10) Wu, H.; Yang, J.; Su, Z. M.; Batten, S. R.; Ma, J. F. J. Am. Chem. Soc. **2011**, 133, 11406.

(11) (a) Carlucci, L.; Ciani, G.; Proserpio, D. M. Coord. Chem. Rev. 2003, 246, 247. (b) Alexandrov, E. V.; Blatov, V. A.; Kochetkov, A. V.; Proserpio, D. M. CrystEngComm 2011, 13, 3947. (c) Delgado-Friedrichs, O.; Foster, M. D.; O'Keeffe, M.; Proserpio, D. M.; Treacy, M. M. J.; Yaghi, O. M. J. Solid State Chem. 2005, 178, 2533. (d) Delgado-Friedrichs, O.; O'Keeffe, M. Acta Crystallogr., Sect. A 2003, 59, 351.

(12) Ke, X. J.; Li, D. S.; Du, M. Inorg. Chem. Commun. 2011, 14, 788 and references cited therein.

(13) Eckhardt, R. H.; Heidl, H.; Fischer, R. D. Chem.—Eur. J. 2003, 9, 1795.

(14) Sava, D. F.; Rohwer, L.; Rodriguez, E. S.; Nenoff, M. A. J. Am. Chem. Soc. 2012, 134, 3983.

(15) Schareina, T.; Schick, C.; Abrahams, B. F.; Kempe, Z. R. Anorg. Allg. Chem. 2001, 627, 1711.

(16) Li, D. S.; Zhang, P.; Zhao, J.; Fang, Z. F.; Du, M.; Zou, K.; Mu, Y. Q. Cryst. Growth Des. **2012**, *12*, 1697.

(17) Carlucci, L.; Ciani, G.; Moret, M.; Proserpio, D. M.; Rizzato, S. Angew. Chem., Int. Ed. **2000**, 39, 1506.

(18) Sun, D.; Ma, S.; Ke, Y.; Collins, D. J.; Zhou, H. C. J. Am. Chem. Soc. 2006, 128, 3896.

(19) Zhao, X. L.; Sun, D.; Yuan, S. A.; Feng, S. Y.; Cao, R.; Yuan, D. Q.; Wang, S. N.; Dou, J. M.; Sun, D. F. *Inorg. Chem.* **2012**, *51*, 10350.

(20) Blatov, V. A. Struct. Chem. 2012, 23, 955; see also http://www.topos.samsu.ru.

(21) Blatov, V. A.; O'Keeffe, M.; Proserpio, D. M. CrystEngComm 2010, 12, 44.

(22) Forgan, R. S.; Sauvage, J.-P.; Stoddart, J. F. Chem. Rev. 2011, 111, 5434.

(23) O'Keeffe, M. Z. Kristallogr. 1991, 196, 21.

(24) O'Keeffe, M.; Peskov, M. A.; Ramsden, S. J.; Yaghi, O. M. Acc.

Chem. Res. 2008, 41, 1782; see also http://rcsr.anu.edu.au/.

(25) Spek, A. L. J. Appl. Crystallogr. 2003, 36, 7.

(26) (a) Maji, T.; Matsuda, K. R.; Kitagawa, S. Nat. Mater. 2007, 6, 142.
(b) Mulfort, K. L.; Farha, O. K.; Malliakas, C. D.; Kanatzidis, M. G.; Hupp, J. T. Chem.—Eur. J. 2010, 16, 276. (c) Yang, S. H.; Lin, X.; Lewis, W.; Suyetin, M.; Bichoutskaia, E.; Parker, J. E.; Tang, C. C.; Allan, D. R.; Rizkallah, P. J.; Hubberstey, P.; Champness, N. R.; Thomas, K. M.; Blake, A. J.; Schröder, S. M. Nat. Mater. 2012, 11, 710.